# Lesson Element

# S-Block elements crossword

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### Across

1. Each Group 1 element has one electron more than the electronic configuration of one of these elements. (5,3)

5. When Group 2 elements are oxidised each atom loses this many electrons. (3)

9. Cations with high charge density can distort or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the cloud of negative charge around a carbonate ion. (8)

12. Within each period Group 1 elements are \_\_\_\_\_\_\_\_ reactive than Group 2 elements. (4)

14. In terms of being reducing agents, Group 1 and 2 metals can be described as this. (6)

15. Common name for the Group 2 elements. (8,5,6)

17. Group 2 oxides form this type of solution with water. (6,8)

18. This Group 1 metal is the least metallic. (7)

20. Collective name given to the elements in Groups 1 and 2 because they have their highest energy electrons in an s sub-shell. (1-5,8)

### Down

2. This compound is more difficult to decompose than strontium carbonate. (6,9)

3. Because they can be cut with a knife, Group 1 metals can be described as this. (4)

4. This property of Group 2 carbonates increases down the group. (7,9)

6. Group 2 elements react vigorously with water. Each element forms the expected ionic what? (5)

7. Group 1 metals have low melting and \_\_\_\_\_\_\_\_\_\_\_\_\_\_ points? (7)

8. Within each group this happens to the level of reactivity as the group descends. (9)

10. Common name for the Group 1 elements. (6,6)

11. Cations at the top of Group 2 have relatively higher charge density because they are \_\_\_\_\_\_\_\_\_\_\_\_\_\_. (7)

13. Mg reacts with water forming the hydroxide and this gas. (8)

16. The hydroxides of Group 1 and Group 2 elements are this. (8)

19. When Group 1 elements are oxidised each atom loses this many electrons. (3)