

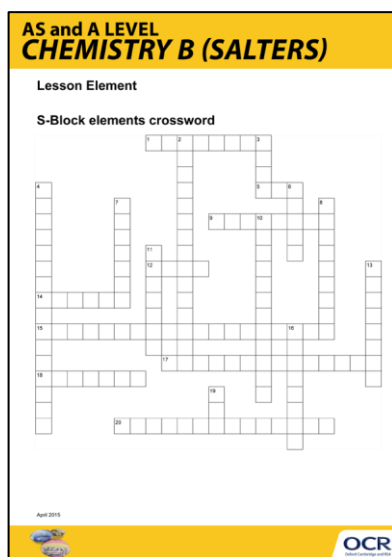
# AS LEVEL CHEMISTRY B (SALTERS)

## Lesson Element

### S-Block elements crossword

#### *Instructions and answers for teachers*

*These instructions should accompany the OCR resource 'S-Block elements crossword' activity which supports OCR AS Level Chemistry B.*



#### Introduction

At the beginning of AS Level, many learners seriously underestimate the rate at which they need to assimilate new terminology into their scientific vocabulary. This crossword is designed to practice or revise some of the key terms related to Group 1 and Group 2 elements, as well as some key ideas associated with this topic.

##### **The Activity:**

This crossword may be used at the end of the topic as a consolidation exercise, or as a revision exercise. Once completed, learners could be asked to make their own glossary or flashcards, choosing ten of the most important words.

##### **Learning outcomes**

This resource relates to the following specification learning outcomes:

EL(o), EL(p), EL(r), EL(u)

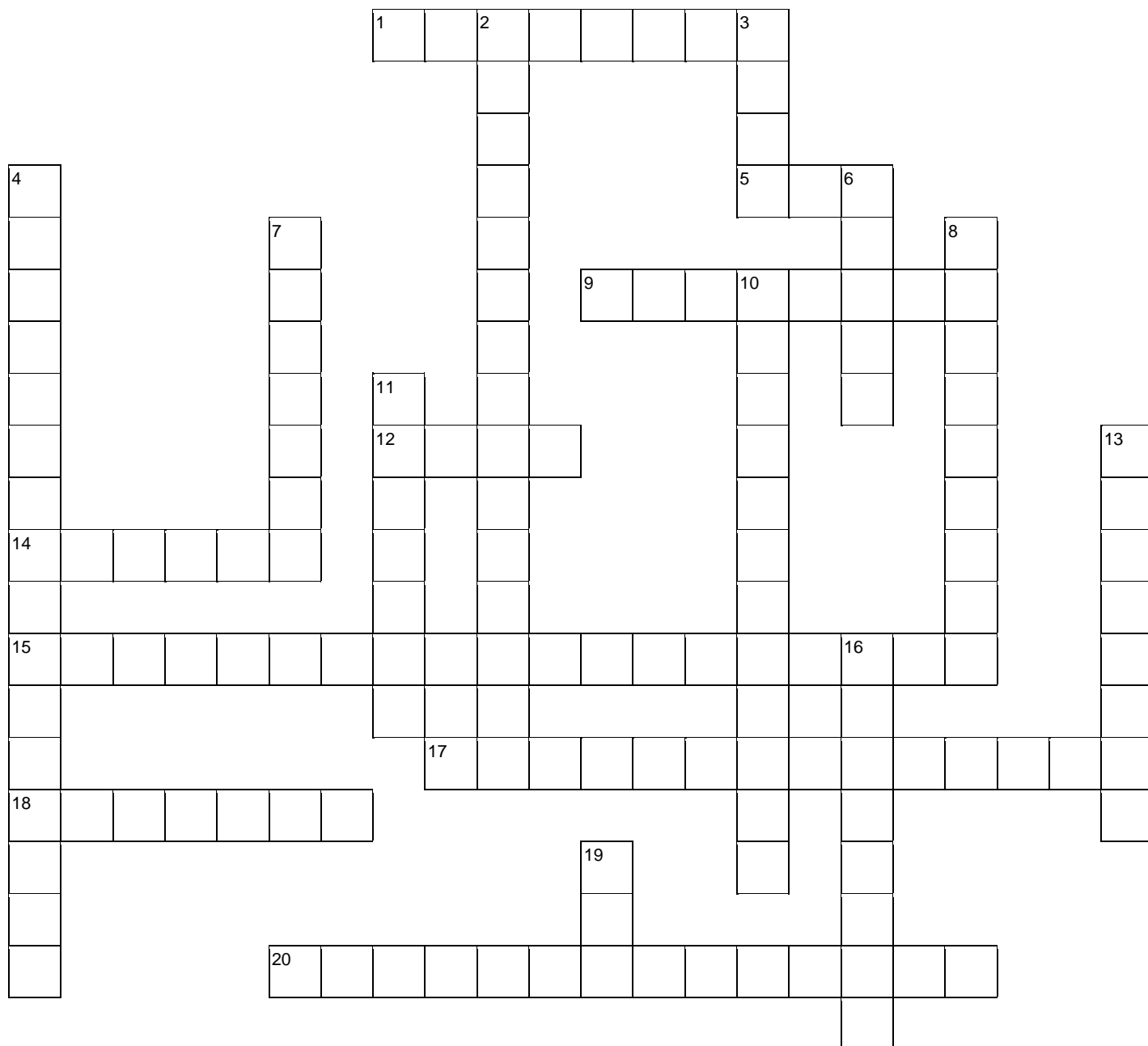
##### **Associated materials:**

'S-Block elements crossword' Lesson Element learner activity sheet.



# AS LEVEL CHEMISTRY B (SALTERS)

## S-Block elements crossword



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## Across

1. Each Group 1 element has one electron more than the electronic configuration of one of these elements. (5,3)
5. When Group 2 elements are oxidised each atom loses this many electrons. (3)
9. Cations with high charge density can distort or \_\_\_\_\_ the cloud of negative charge around a carbonate ion. (8)
12. Within each period Group 1 elements are \_\_\_\_\_ reactive than Group 2 elements. (4)
14. In terms of being reducing agents, Group 1 and 2 metals can be described as this. (6)
15. Common name for the Group 2 elements. (8,5,6)
17. Group 2 oxides form this type of solution with water. (6,8)
18. This Group 1 metal is the least metallic. (7)
20. Collective name given to the elements in Groups 1 and 2 because they have their highest energy electrons in an s sub-shell. (1-5,8)

## Down

2. This compound is more difficult to decompose than strontium carbonate. (6,9)
3. Because they can be cut with a knife, Group 1 metals can be described as this. (4)
4. This property of Group 2 carbonates increases down the group. (7,9)
6. Group 2 elements react vigorously with water. Each element forms the expected ionic what? (5)
7. Group 1 metals have low melting and \_\_\_\_\_ points? (7)
8. Within each group this happens to the level of reactivity as the group descends. (9)
10. Common name for the Group 1 elements. (6,6)
11. Cations at the top of Group 2 have relatively higher charge density because they are \_\_\_\_\_. (7)
13. Mg reacts with water forming the hydroxide and this gas. (8)
16. The hydroxides of Group 1 and Group 2 elements are this. (8)
19. When Group 1 elements are oxidised each atom loses this many electrons. (3)



# AS LEVEL CHEMISTRY B (SALTERS)

## Answers

### Across

1. noble gas
5. two
9. polarise
12. more
14. strong
15. alkaline earth metals
17. weakly alkaline
18. lithium
20. s-block elements

### Down

2. barium carbonate
3. soft
4. thermal stability
6. oxide
7. boiling
8. decreases
10. alkali metals
11. smaller
13. hydrogen
16. alkaline
19. one



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April 2015

