Modified Enlarged 24pt

OXFORD CAMBRIDGE AND RSA EXAMINATIONS

Monday 20 June 2022 – Morning

GCSE (9–1) Chemistry B (Twenty First Century Science)

J258/04 Depth in Chemistry (Higher Tier)

Time allowed: 1 hour 45 minutes plus your additional time allowance

YOU MUST HAVE:
a ruler (cm/mm)
the Data Sheet for GCSE (9–1) Chemistry B (with this document)
Loose Sheet for Question 1

YOU CAN USE:
a scientific or graphical calculator
an HB pencil

Please write clearly in black ink.

Centre number | | | | | |

Candidate number | | | | |

First name(s) _____

Last name _____

READ INSTRUCTIONS OVERLEAF

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INSTRUCTIONS

Use black ink. You can use an HB pencil, but only for graphs and diagrams.

Write your answer to each question in the space provided. If you need extra space use the lined pages at the end of this booklet. The question numbers must be clearly shown.

Answer ALL the questions.

Where appropriate, your answer should be supported with working. Marks might be given for using a correct method, even if your answer is wrong.

INFORMATION

The total mark for this paper is 90.

The marks for each question are shown in brackets [ ].

Quality of extended response will be assessed in questions marked with an asterisk (*).

ADVICE

Read each question carefully before you start your answer.

Answer ALL the questions.

1 FIG. 1.1 on the Loose Sheet shows the uses, properties and structure of some substances which contain carbon atoms.

(a) Petrol contains cyclo-octane.

(i) Which symbol should be used to warn people of the main hazard when handling cyclo-octane?

Tick (✓) ONE box. [1]

☐

☐

☐

☐

(ii) Write down TWO safety precautions people should take when filling their cars with petrol.

1 _____

_____

2 _____

_____

[2]

(b) Carbon makes a much bigger range of different types of molecules than any other element.

Give ONE reason why carbon atoms can form so many different types of molecules.

Use FIG. 1.1 to help you to answer.

_____

_____ [1]

(c) **Explain why polyethene and synthetic rubber have different properties.**

Use ideas about structure from FIG. 1.1 to help you to answer.

_____

_____

_____

_____ **[2]**

(d) (i) How are nanoparticles different to the other substances in FIG. 1.1?

_____

_____ [1]

(ii) Explain why the structure of carbon nanoparticles helps them to carry medicine molecules into the body.

_____

_____ [2]

2 **Ammonia is used to make fertilisers. It is produced in a large-scale process.**

The equation shows the reaction that happens in the process:

$$N_2 + 3H_2 \rightleftharpoons \underset{\text{ammonia}}{2NH_3}$$

The process happens in a reactor.

(a) The percentage yield of ammonia is usually between 10% and 20%.

(i) Use the equation to explain why it is not possible to get 100% yield of ammonia.

_____

_____ [1]

(ii) The gas that leaves the reactor contains ammonia mixed with two other gases.

Use the equation to help you to give the names of the other TWO gases.

_____ AND _____ [1]

(b) An ammonia factory tests a new reactor.

The table shows the theoretical yield and actual yield for a process in the new reactor.

Theoretical yield (tonnes)	150
Actual yield (tonnes)	19.5

Calculate the percentage yield for the process in the new reactor.

Use the equation:

$$\text{percentage yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100\%$$

percentage yield = _____ % [2]

(c) Ammonia is very soluble in water.

Kofi does an experiment to make some ammonia.

He tries to collect it using the apparatus shown.

(i) Bubbles of ammonia gas enter the water but no gas collects in the test-tube.

What happens to the ammonia gas when it enters the water?

_____

_____ [1]

(ii) **Which apparatus should Kofi use to collect ammonia?**

Tick (✓) ONE box. [1]

- burette ☐
- gas syringe ☐
- measuring cylinder ☐
- volumetric flask ☐

(d) Ammonia is an alkaline gas.

The pH of ammonia solution can be measured using a pH meter.

(i) Predict the pH of ammonia.

pH _____ [1]

(ii) Describe another method you could use to measure the pH of ammonia solution.

_____

_____

_____ [2]

3 The graph shows the change in carbon dioxide concentration in the Earth's atmosphere over the last 575 million years.

(a) (i) How many millions of years ago was the concentration of carbon dioxide at its highest point?

_____ million years ago [1]

(ii) What was the concentration of carbon dioxide at its highest point?

_____ ppm (parts per million) [1]

Carbon dioxide concentration in ppm (parts per million)
8000
7000
6000
5000
4000
3000
2000
1000
0
Current CO_2 concentration
600 550 500 450 400 350 300 250 200 150 100 50 0
Millions of years ago

(b) Ben looks at the graph and says, ‘Carbon dioxide concentration was always much higher in the past than it is today.’

Does the graph support this statement?

Tick (✓) ONE box.

Yes ☐

No ☐

Use values from the graph to explain your answer.

_____

_____

_____ [2]

(c) Scientists say that the concentration of carbon dioxide is now 0.04 %.

In the year 2000 the concentration of carbon dioxide was 370 ppm (parts per million).

Do a calculation to find out the difference between these two values.

Give your answer in ppm.

Use this formula:

$$\text{concentration in ppm} = \text{concentration in \%} \times 10^4$$

difference between values _____ ppm [2]

(d) **Some scientists fear that increased carbon dioxide in the air may lead to global food shortages.**

Explain why this could happen.

_____

_____ **[1]**

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4 The table shows information about some alkanes.

(a) The empirical formula of octane is C_4H_9.

(i) What is the EMPIRICAL formula of pentacontane?

_____ **[1]**

(ii) The empirical formulae of methane and propane are the same as their molecular formulae.

Explain why.

_____

_____ **[1]**

Alkane	Molecular formula	Relative formula mass	Melting point (°C)	State at room temperature (at 20 °C)
methane	CH_4	16	–182	gas
propane	C_3H_8	44	–190	gas
octane	C_8H_{18}	114	–57	liquid
pentacontane	$C_{50}H_{102}$	702	93	solid

(b) Predict the BOILING point of octane.

Explain your reasoning.

Boiling point _____°C

Reason _____

_____ [2]

(c)* A text book has this description of a homologous series.

'A homologous series contains compounds with similar structures, the same general formula and similar chemical properties. The compounds show a trend in physical properties down the series.'

Use your own knowledge and the information in the table to explain how the alkanes match this description of a homologous series. [6]

_____

_____

5 Some metals are extracted from their ores by heating with carbon.

Other metals are extracted from their ores by electrolysis.

FIG. 5.1 shows the order of reactivity of some metals, compared to carbon.

FIG. 5.1

most reactive	sodium	extracted by electrolysis
	calcium	
	magnesium	
	[carbon]	
	zinc	extracted by heating with carbon
least reactive	copper	

(a) Explain why calcium and copper must be extracted by different methods.

_____

_____ [2]

(b) What are the most common methods of extraction for iron, potassium, aluminium and lead from their ores?

Use ideas about the order of reactivity of the metals to help you to answer.

Put a tick (✓) in ONE box in each row. [2]

Metal	Extracted by heating with carbon	Extracted by electrolysis
iron		
potassium		
aluminium		
lead		

(c) Some copper ores contain copper sulfide.

Copper is extracted from copper sulfide in a two-stage process.

STAGE 1: Copper sulfide reacts with oxygen:

$$Cu_2S(s) + 2O_2(g) \longrightarrow 2CuO(s) + SO_2(g)$$

STAGE 2: Copper oxide reacts with carbon:

$$2CuO(s) + C(s) \longrightarrow 2Cu(s) + CO_2(g)$$

(i) Look at the equation for STAGE 1.

Predict whether the mass of the solid increases, decreases or stays the same during STAGE 1.

Use the Periodic Table and calculations of relative formula masses to help you.

Prediction _____ [3]

(ii) During STAGE 2, the total mass of the solids decreases.

Explain why STAGE 2 obeys the law of conservation of mass even though the total mass of the solids decreases.

_____

_____

_____ [2]

(iii) The reaction of copper in STAGE 2 can be represented by a half equation:

$$Cu^{2+} + 2e^{-} \longrightarrow Cu$$

Is copper oxidised or reduced in this half equation?

Explain your answer.

Oxidised or reduced? _____

Explanation _____

_____ [1]

6 FIG. 6.1 shows the particle model for solids, liquids and gases.

FIG. 6.1

(a) Neon is a gas in Group 18 (0) of the Periodic Table.

Use the particle model in FIG. 6.1 to describe what happens to the arrangement and movement of the atoms when neon changes from a liquid to a gas.

_____

_____

_____ [2]

(b) **Sodium atoms form sodium ions with a single positive charge.**

Neon atoms do not form ions.

Use ideas about electrons to explain these statements.

_____

_____

_____

_____ **[3]**

(c)* Sodium chloride contains sodium ions, Na^+, and chloride ions, Cl^-.

FIG. 6.2 shows the ionic model for solid and liquid sodium chloride.

FIG. 6.2

The melting points of neon and sodium chloride are shown in the table.

	Melting point (°C)
neon	–249
sodium chloride	801

Describe the similarities and differences in how melting is represented by the models in FIG. 6.1 AND FIG. 6.2 and explain why both models are needed to explain the differences in the melting points shown in the table.

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____

_____ **[6]**

7 Jane does an experiment.

She puts 10 cm^3 of copper sulfate solution in a boiling tube. She adds 5 cm^3 of sodium hydroxide solution. A precipitate of copper hydroxide forms.

(a) Complete the WORD and SYMBOL equations for the reaction in the boiling tube by filling in the name of the missing product and the state symbols for each substance. [3]

copper sulfate + sodium hydroxide → copper hydroxide + _____

$CuSO_4$(___) + 2NaOH(___) → $Cu(OH)_2$(___) + Na_2SO_4(___)

(b) The final mixture contains a precipitate of copper hydroxide mixed with a solution of other dissolved substances.

Jane wants to separate pure copper hydroxide from this mixture.
She wants to make sure that she removes any traces of other dissolved substances from the precipitate.

(i) Describe how she can separate pure copper hydroxide from the final mixture.

_____

_____

_____

_____ [2]

(ii) Jane leaves the copper hydroxide to dry in a warm oven. After 30 minutes she weighs the copper hydroxide on a balance.

Jane is not sure if the copper hydroxide is completely dry.

Suggest how Jane can use the oven and the balance to show that the copper hydroxide is completely dry.

_____

_____

_____ [2]

(c) Jane does more experiments.

She adds a different volume of sodium hydroxide solution to 20 cm^3 of copper sulfate solution each time.

She records the mass of dry copper hydroxide that forms in each experiment.

The table shows her results.

Volume of copper sulfate solution (cm^3)	Volume of sodium hydroxide solution added (cm^3)	Mass of dry copper hydroxide formed (g)
20	5	0.25
20	10	0.49
20	15	0.75
20	20	0.98
20	25	0.98
20	30	0.98

(i) Plot Jane's results on the graph.

Draw lines of best fit to show the pattern in the results. [3]

1.00
0.80
0.60
0.40
0.20
0.00
Mass of dry copper hydroxide formed (g)
0
5
10
15
20
25
30
35
Volume of sodium hydroxide solution added (cm3)

(ii) Suggest why the mass of the dry precipitate does not continue to increase when more than 20 cm^3 of sodium hydroxide solution is added.

_____

_____ [1]

(iii) Jane and Alex look at the table and the equation for the reaction:

$$CuSO_4 + 2NaOH \longrightarrow Cu(OH)_2 + Na_2SO_4$$

They disagree about the results.

Jane says, 'I think the concentration of copper sulfate solution is the same as the concentration of sodium hydroxide solution.'

Alex says, 'I think the sodium hydroxide solution is double the concentration of the copper sulfate solution.'

Who is right?

Jane ☐

Alex ☐

Use information from the table and the equation to explain your choice.

_____

_____

_____

_____ **[2]**

8 **Beth works for a company that makes hydrogen fuel cells.**

She measures the potential difference of the cell shown.

(a) **Before she sets up the cell, Beth tests each gas to check its identity.**

Describe the tests and the results for hydrogen and oxygen gas.

hydrogen test _____

result _____

oxygen test _____

result _____

[2]

(b) These half equations show the reactions that happen at each electrode in the fuel cell:

at the hydrogen electrode:

$$2H_2 + 4OH^- \longrightarrow 4H_2O + 4e^-$$

at the oxygen electrode:

$$O_2 + 2H_2O + 4e^- \longrightarrow 4OH^-$$

(i) Beth wants to make sure that she gets the highest possible potential difference from the fuel cell.

She makes sure that she uses double the volume of hydrogen compared to oxygen.

Use the half equations to explain why she needs to do this.

_____

_____ [1]

(ii) Use the half equations to write an overall equation for the reaction that happens in the fuel cell.

_____ [2]

(iii) A fuel cell filled with potassium hydroxide solution works better than a fuel cell filled with pure water.

One reason it works better is because potassium hydroxide solution is a better electrical conductor than pure water.

Explain why potassium hydroxide solution is a better electrical conductor than pure water.

_____

_____ [1]

(iv) Use the equations to suggest one other reason why using potassium hydroxide solution helps the fuel cell to work better.

_____

_____ [1]

(c) Beth uses a fixed amount of hydrogen and oxygen gas. After a time the potential difference of the cell decreases.

Explain why this happens.

_____

_____ [1]

(d) Beth's company want to use the cell to provide power for a submarine.

Submarines travel deep under the surface of the sea.

Most submarines have engines that burn diesel fuel.

Beth thinks that submarines that burn diesel fuel produce waste that is much more harmful to the sea than submarines that use hydrogen fuel cells.

(i) Explain why she is correct.

_____

_____

_____ [2]

(ii) Give one DISADVANTAGE of using a hydrogen fuel cell rather than diesel for a submarine.

_____

_____ [1]

9 The graph shows the solubility of some Group 1 salts in water.

(a) (i) Use data from the graph to calculate the solubility of sodium chloride at 50 °C.

Use the information:
100 g of water = 100 cm^3 of water.

Give your answer in g/dm^3.

Solubility of
sodium chloride at 50 °C _____ g/dm^3 [2]

(ii) Calculate the concentration of sodium chloride at 50 °C in mol / dm^3.

Use the equations:

$$\text{number of moles} = \frac{\text{mass of substance (g)}}{\text{relative formula mass (g)}}$$

and

$$\text{concentration (mol / dm}^3\text{)} = \frac{\text{number of moles of solute}}{\text{volume (dm}^3\text{)}}$$

Use your answer to (a)(i).

Give your answer to 3 significant figures.

Concentration of sodium chloride at 50 °C _____ mol / dm^3 [3]

(b) Kai looks at the graph and writes down a mathematical relationship.

(i) Kai writes:

solubility of potassium chloride α temperature

Use the graph to explain in words why this relationship is INCORRECT.

_____

_____

_____ [2]

(ii) Use data from the graph to do a calculation to predict the solubility of potassium chloride at 100 °C.

Solubility _____ g in 100 g of water [3]

END OF QUESTION PAPER

ADDITIONAL ANSWER SPACE

If additional space is required, you should use the following lined page(s). The question number(s) must be clearly shown in the margin(s).

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