

GENERAL CERTIFICATE OF SECONDARY EDUCATION

TWENTY FIRST CENTURY SCIENCE

A172/02

CHEMISTRY A

Unit A172: Modules C4, C5, C6 (Higher Tier)

Candidates answer on the question paper
 A calculator may be used for this paper

OCR Supplied Materials:

None

Duration: 1 hour

Other Materials Required:

- Pencil
- Ruler (cm/mm)

Candidate Forename		Candidate Surname	
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Centre Number						Candidate Number				
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INSTRUCTIONS TO CANDIDATES

- Write your name clearly in capital letters, your Centre Number and Candidate Number in the boxes above.
- Use black ink. Pencil may be used for graphs and diagrams only.
- Read each question carefully and make sure that you know what you have to do before starting your answer.
- Answer **all** the questions.
- Write your answer to each question in the space provided, however additional paper may be used if necessary.

INFORMATION FOR CANDIDATES

- Your quality of written communication is assessed in questions marked with a pencil (✎).
- A list of qualitative tests for ions is printed on page 2.
- The Periodic Table can be found on the back page.
- The number of marks for each question is given in brackets [] at the end of the question or part question.
- The total number of marks for this paper is **60**.
- This document consists of **24** pages. Any blank pages are indicated.

For Examiner's Use		
	Max	Mark
1	3	
2	2	
3	6	
4	3	
5	4	
6	3	
7	2	
8	6	
9	6	
10	3	
11	2	
12	2	
13	2	
14	12	
15	4	
TOTAL	60	

TWENTY FIRST CENTURY SCIENCE DATA SHEET

Qualitative analysis

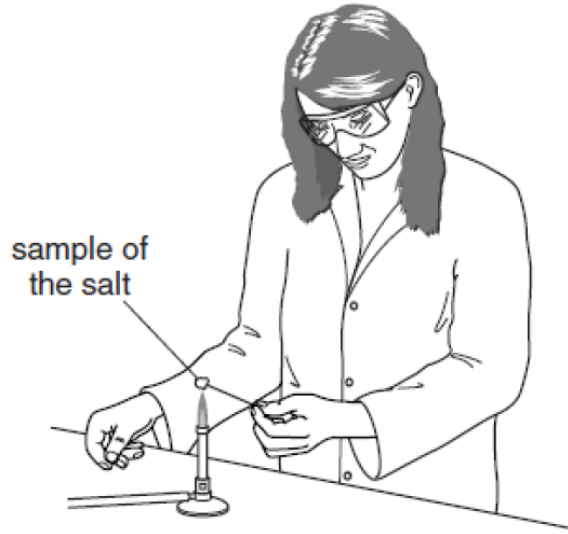
Tests for ions with a positive charge

Ion	Test	Observation
calcium Ca^{2+}	add dilute sodium hydroxide	a white precipitate forms; the precipitate does not dissolve in excess sodium hydroxide
copper Cu^{2+}	add dilute sodium hydroxide	a light blue precipitate forms; the precipitate does not dissolve in excess sodium hydroxide
iron(II) Fe^{2+}	add dilute sodium hydroxide	a green precipitate forms; the precipitate does not dissolve in excess sodium hydroxide
iron(III) Fe^{3+}	add dilute sodium hydroxide	a red-brown precipitate forms; the precipitate does not dissolve in excess sodium hydroxide
zinc Zn^{2+}	add dilute sodium hydroxide	a white precipitate forms; the precipitate dissolves in excess sodium hydroxide

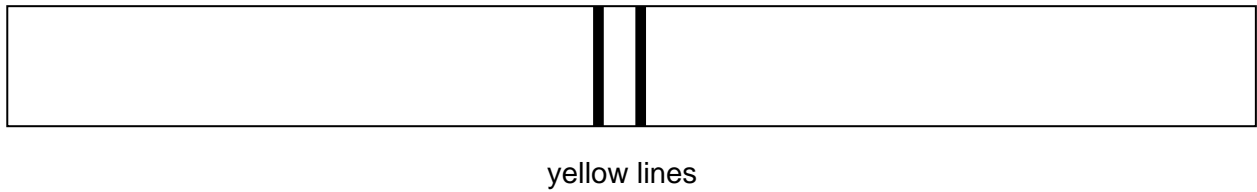
Tests for ions with a negative charge

Ion	Test	Observation
carbonate CO_3^{2-}	add dilute acid	the solution effervesces; carbon dioxide gas is produced (the gas turns lime water from colourless to milky)
chloride Cl^-	add dilute nitric acid, then add silver nitrate	a white precipitate forms
bromide Br^-	add dilute nitric acid, then add silver nitrate	a cream precipitate forms
iodide I^-	add dilute nitric acid, then add silver nitrate	a yellow precipitate forms
sulfate SO_4^{2-}	add dilute acid, then add barium chloride or barium nitrate	a white precipitate forms

2 Eve tests some salts by doing a flame test.



Eve heats a sodium salt. She sees that it gives off a coloured light.
She looks at the spectrum of light through a spectroscope.
She sees some yellow lines.



Eve then heats a potassium salt and looks at the spectrum of light it gives off.
Write down one **similarity** and one **difference** between the two spectra that Eve sees.

.....

.....

.....

..... [2]

[Total: 2]

4 Sodium reacts with the halogens.

The reaction takes place between sodium metal and the halogen in the gas state.

- (a) The table shows what happens when hot sodium is put into jars containing different halogen gases.

halogen gas	appearance of halogen gas at start	time for reaction to finish in seconds	appearance of product at end
chlorine	pale green	5	white solid
bromine			
iodine	purple	15	white solid

Complete the table to describe what you would see when sodium is put into a jar containing bromine gas.

[2]

- (b) What is the name and formula of the white solid left at the end of the reaction when sodium reacts with iodine?

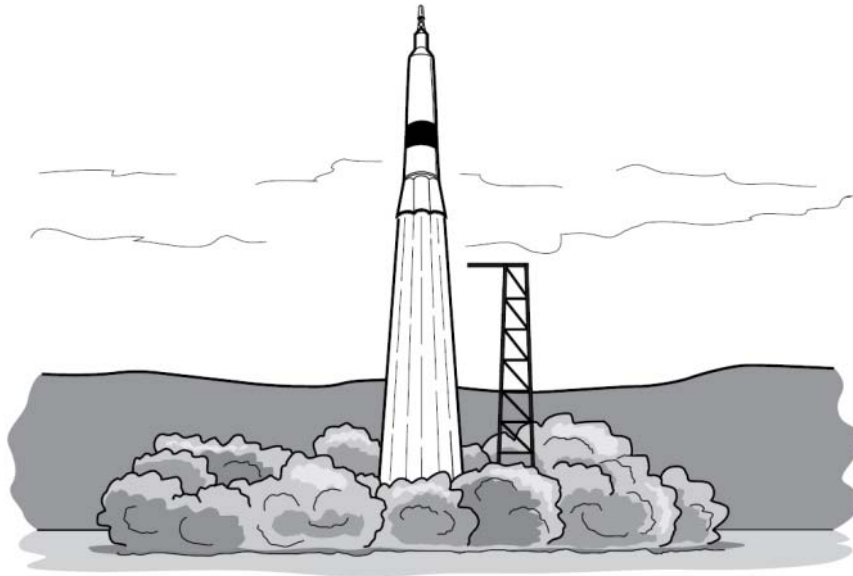
.....

..... [1]

[Total: 3]

5 Lithium is an element in Group 1 of the Periodic Table.

It can be added to rocket fuel to give an extra boost for take off.



(a) Lithium works well in rocket fuels because it is very reactive.

Which of the following statements about the reactivity of lithium are **true** and which are **false**?

Put a tick (✓) in the correct box in each row.

	true	false
Lithium reacts with cold water.		
Lithium reacts with other Group 1 elements to form compounds.		
Lithium tarnishes in moist air more quickly than potassium.		
Lithium chloride is very unstable.		

[2]

(b) When the fuel burns, the lithium also burns.

Complete the balanced symbol equation to show what happens when lithium burns.

word equation lithium + oxygen → lithium oxide

balanced symbol equation + →Li₂O

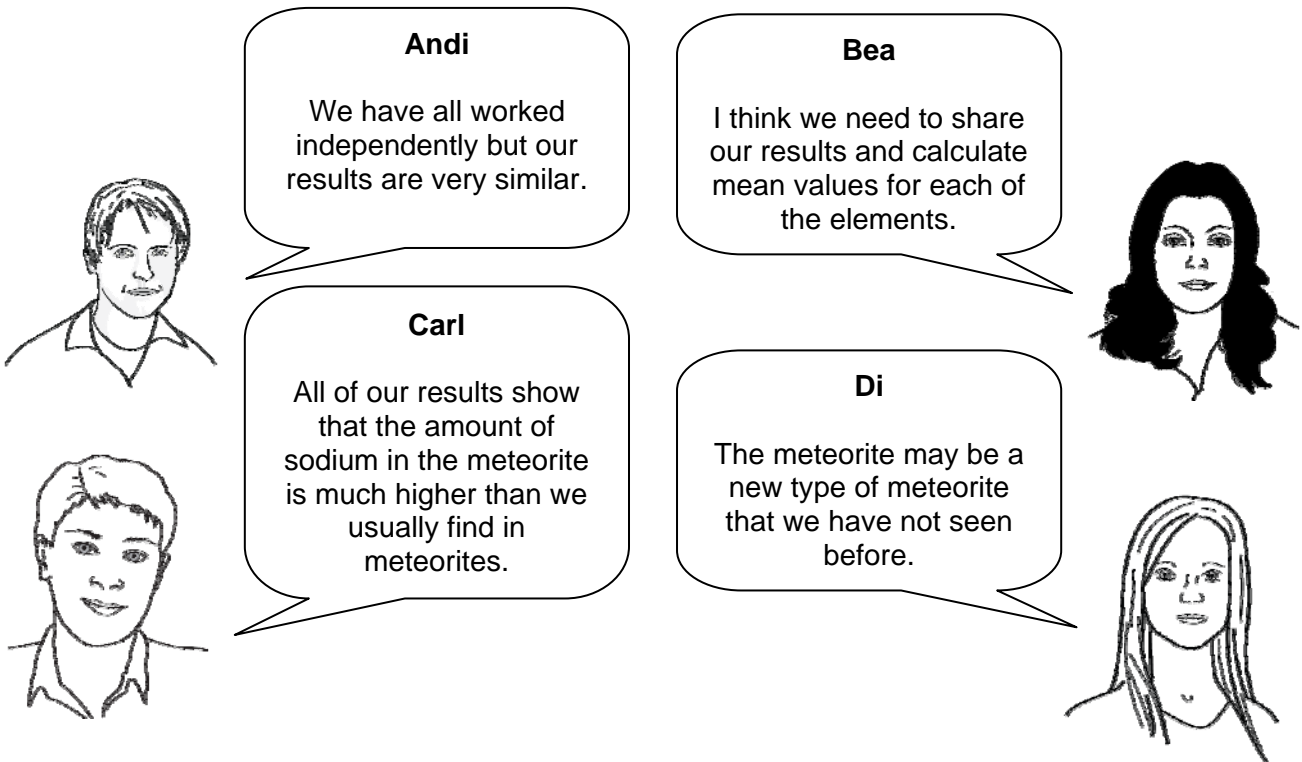
[2]

[Total: 4]

6 Meteorites are small pieces of rock that land on Earth from space.

Four scientists are investigating the amounts of different elements in a meteorite.

They talk about what they find.



Andi
We have all worked independently but our results are very similar.

Bea
I think we need to share our results and calculate mean values for each of the elements.

Carl
All of our results show that the amount of sodium in the meteorite is much higher than we usually find in meteorites.

Di
The meteorite may be a new type of meteorite that we have not seen before.

(a) Put ticks (✓) in the table to identify who is making each type of statement.

	Andi	Bea	Carl	Di	none of the scientists
Who is talking about the reproducibility of data?					
Who is suggesting a hypothesis?					
Who is talking about an outlier?					
Who is making a suggestion that would help to work out a best estimate?					

[2]

(b) Di thinks that the meteorite is a new type of meteorite.

She wants to convince other scientists.

Here are some actions that she could take.

- A publish a report in a newspaper so that everyone can see it
- B publish a report in a scientific journal so that other scientists can see it

- C repeat the experiment to confirm her results before she tells anyone else
- D tell everyone immediately because it is important

- E ask other scientists to repeat the experiment
- F ask other scientists to identify outliers in her data

Choose one action from each pair, then write the letters in the boxes to show the order in which she should do them.

start

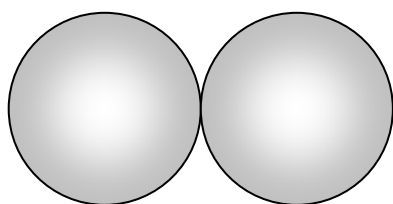
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 end

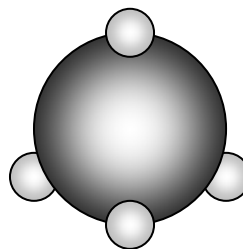
[1]

[Total: 3]

- 7 These diagrams show the arrangement of atoms in a molecule of oxygen and a molecule of methane.



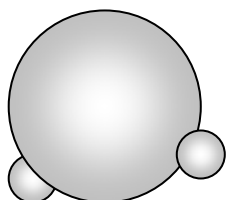
oxygen



methane

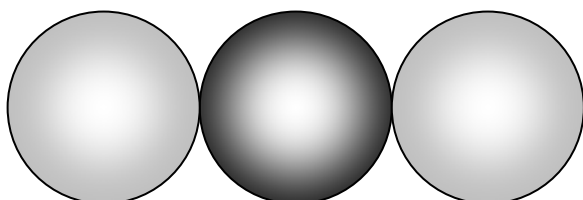


What is the name and formula for each of the following molecules?



name

formula



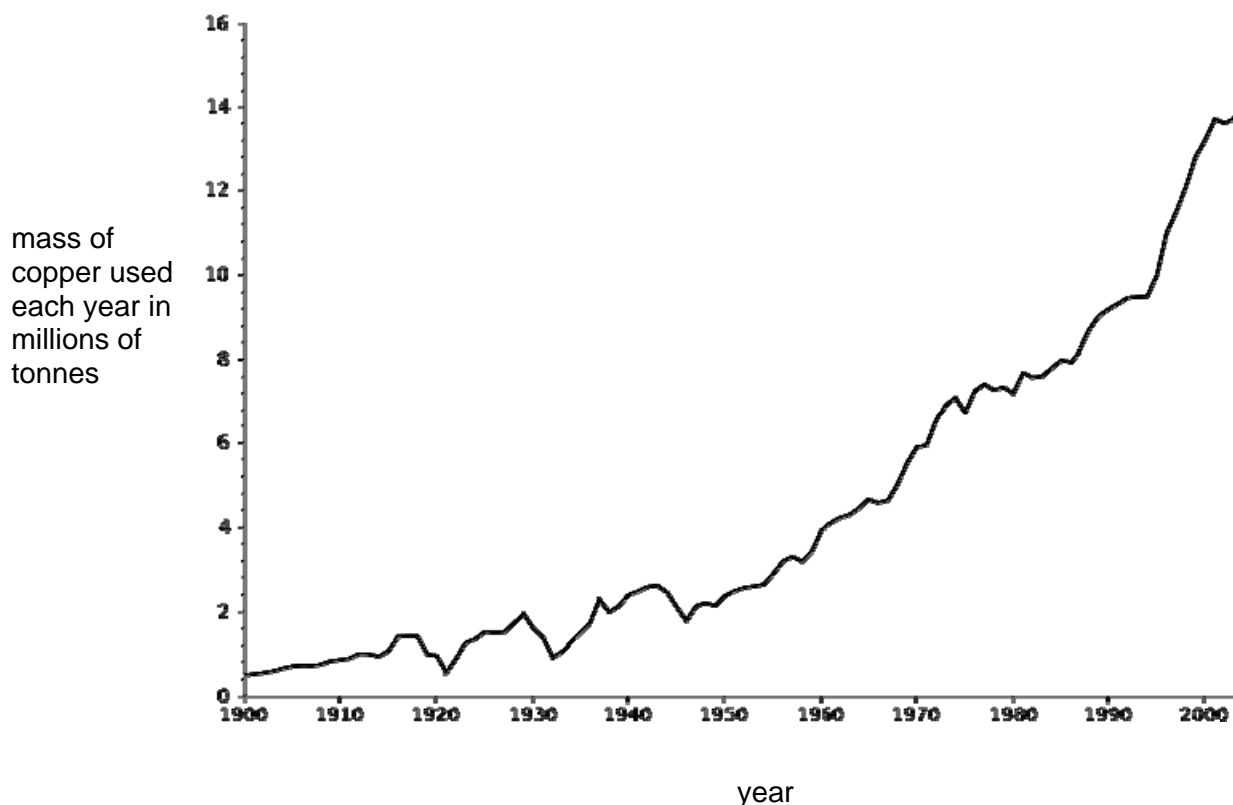
name

formula

[2]

[Total: 2]

9 The graph shows the mass of copper used each year during the last century.



Copper is a fairly common metal. It accounts for approximately 0.007 % of the Earth's crust.

(a) (i) Use your knowledge of copper extraction and the graph to predict whether we will be able to produce enough copper to meet demand in the future.

Explain your answer.

.....

.....

.....

..... [2]

(ii) Suggest how it might be possible to reduce the mass of copper that we need to extract each year.

.....

..... [1]

(b) Copper is obtained by mining copper ores from the ground.

A mining company can recover 40 g of copper ore from every kilogram of mined rock.

The ore they recover is called cuprite, and has the formula Cu_2O .

(i) What is the maximum amount of copper **metal** the company will be able to recover from a kilogram of mined rock?

answer = g [2]

(ii) Complete the sentences to explain how copper metal can be extracted from cuprite.

Cuprite ore is with carbon.

Copper is during the reaction because

it loses

Carbon is during the reaction because

it gains

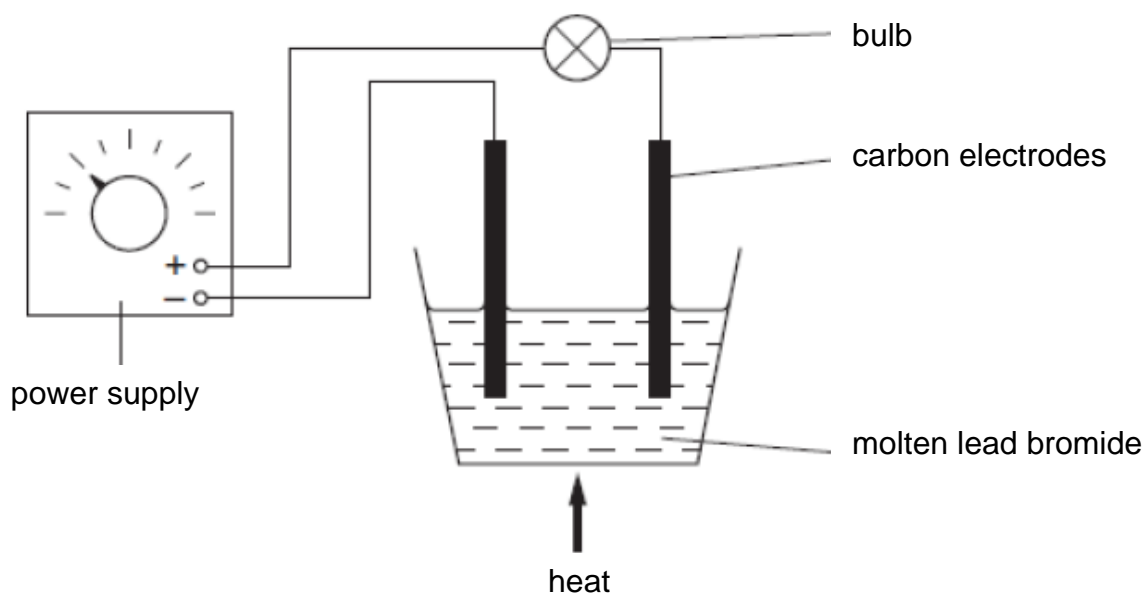
[1]

[Total: 6]

10 John does an experiment.

He passes electricity through molten lead bromide.

The diagram shows how he sets up his experiment.



(a) Lead bromide is an ionic compound.

What will form at each electrode when John turns on the power supply?

.....

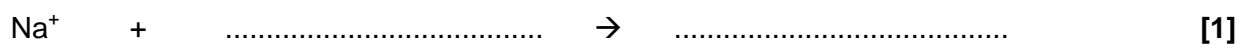
.....

.....

..... [2]

(b) John finds out that atoms of sodium metal can be made from sodium chloride by electrolysis of molten sodium chloride.

Complete the equation to show what happens when a sodium ion forms a sodium atom.



[Total: 3]

11 Silicon dioxide is a giant molecular compound.

The atoms in silicon dioxide are held together by covalent bonds.

Which statements about the bonding in silicon dioxide are correct?

Put ticks (✓) in the boxes next to the **two** correct answers.

Electrons are gained or lost to form a full outer shell.

Electrons are shared between atoms.

The nucleus of each bonded atom attracts electrons.

Charged ions are attracted towards one another.

The nuclei of the atoms attract each other.

[2]

[Total: 2]

12 This question is about solids and liquids.

(a) Which is the solid acid in this list?

Put a **ring** around the correct answer.

ethanoic acid

nitric acid

sulfuric acid

citric acid

[1]

(b) Baking powder contains small grains of a solid acid and small grains of a solid alkali.

The acid in baking powder does not react with the alkali until water is added.

Explain why the reaction only starts when water is added.

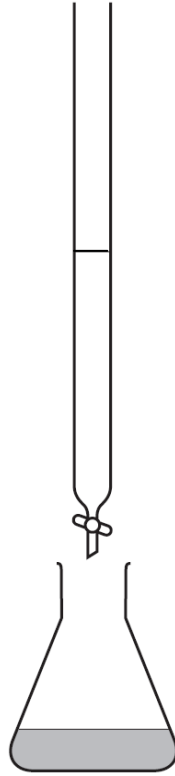
.....

.....

..... [1]

[Total: 2]

13 Mary does a titration.



She puts 25.0 cm^3 of alkali solution in a conical flask. She adds a few drops of indicator to the alkali and then adds acid from the burette.

She does a rough titration first. She then does an accurate titration.

Describe one thing Mary should do to make her second titration as accurate as possible, and explain why this increases the accuracy.

.....

.....

.....

.....

.....

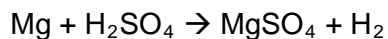
.....

.....

[2]

[Total: 2]

- 14 Bobby reacts 3 g of magnesium pieces with an excess of sulfuric acid until all of the magnesium has reacted.



- (a) What mass of magnesium sulfate will be produced by the reaction?

Show your working.

mass of magnesium sulfate = g [2]

- (b) Bobby collects the hydrogen gas produced by the reaction. Every 10 seconds he records the volume of gas that has been collected.

Here are his results.

time after start of reaction in s	volume of gas collected in cm ³
0	0
10	20
20	30
30	35
40	35
50	35

What was the rate of reaction during the first 10 seconds?

Show your working and include appropriate units in your answer.

rate of reaction = [1]

(c) Bobby does the experiment a further four times.

Each time he makes **one** change to the way he does the experiment.

experiment	volume of gas collected after 10s, in cm ³	volume of gas collected after 30s, in cm ³	volume of gas collected after 50s, in cm ³
original experiment	20	35	35
experiment A	35	40	40
experiment B	30	35	35
experiment C	20	30	35
experiment D	25	35	35

In which experiment did Bobby use a larger mass of magnesium pieces?

Explain your answer.

.....

.....

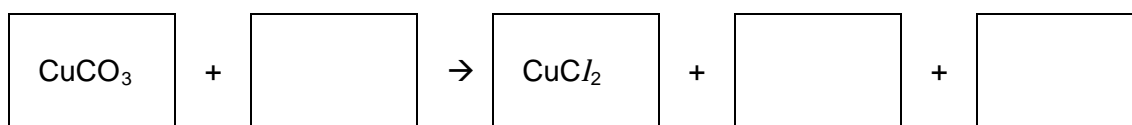
.....

..... [3]

15 Geoff reacts copper carbonate with hydrochloric acid.

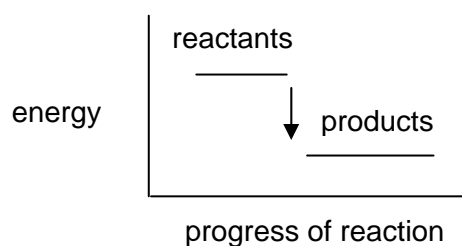
He knows that this will produce a salt and two other products.

(a) Write the formula of each chemical in the correct box, then balance the equation.



[3]

(b) Geoff draws an energy level diagram for the reaction.



What is the name given to this type of reaction?

..... [1]

[Total: 4]

[Paper Total: 60]

END OF QUESTION PAPER

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Periodic Table

1	2											3	4	5	6	7	0											
		<div style="border: 1px solid black; padding: 5px; display: inline-block;"> Key relative atomic mass atomic symbol <small>name</small> atomic (proton) number </div>																									<div style="border: 1px solid black; padding: 5px; display: inline-block;"> 1 H <small>hydrogen</small> 1 </div>	<div style="border: 1px solid black; padding: 5px; display: inline-block;"> 4 He <small>helium</small> 2 </div>
7 Li <small>lithium</small> 3	9 Be <small>beryllium</small> 4											11 B <small>boron</small> 5	12 C <small>carbon</small> 6	14 N <small>nitrogen</small> 7	16 O <small>oxygen</small> 8	19 F <small>fluorine</small> 9	20 Ne <small>neon</small> 10											
23 Na <small>sodium</small> 11	24 Mg <small>magnesium</small> 12											27 Al <small>aluminium</small> 13	28 Si <small>silicon</small> 14	31 P <small>phosphorus</small> 15	32 S <small>sulfur</small> 16	35.5 Cl <small>chlorine</small> 17	40 Ar <small>argon</small> 18											
39 K <small>potassium</small> 19	40 Ca <small>calcium</small> 20	45 Sc <small>scandium</small> 21	48 Ti <small>titanium</small> 22	51 V <small>vanadium</small> 23	52 Cr <small>chromium</small> 24	55 Mn <small>manganese</small> 25	56 Fe <small>iron</small> 26	59 Co <small>cobalt</small> 27	59 Ni <small>nickel</small> 28	63.5 Cu <small>copper</small> 29	65 Zn <small>zinc</small> 30	70 Ga <small>gallium</small> 31	73 Ge <small>germanium</small> 32	75 As <small>arsenic</small> 33	79 Se <small>selenium</small> 34	80 Br <small>bromine</small> 35	84 Kr <small>krypton</small> 36											
85 Rb <small>rubidium</small> 37	88 Sr <small>strontium</small> 38	89 Y <small>yttrium</small> 39	91 Zr <small>zirconium</small> 40	93 Nb <small>niobium</small> 41	96 Mo <small>molybdenum</small> 42	[98] Tc <small>technetium</small> 43	101 Ru <small>ruthenium</small> 44	103 Rh <small>rhodium</small> 45	106 Pd <small>palladium</small> 46	108 Ag <small>silver</small> 47	112 Cd <small>cadmium</small> 48	115 In <small>indium</small> 49	119 Sn <small>tin</small> 50	122 Sb <small>antimony</small> 51	128 Te <small>tellurium</small> 52	127 I <small>iodine</small> 53	131 Xe <small>xenon</small> 54											
133 Cs <small>caesium</small> 55	137 Ba <small>barium</small> 56	139 La* <small>lanthanum</small> 57	178 Hf <small>hafnium</small> 72	181 Ta <small>tantalum</small> 73	184 W <small>tungsten</small> 74	186 Re <small>rhenium</small> 75	190 Os <small>osmium</small> 76	192 Ir <small>iridium</small> 77	195 Pt <small>platinum</small> 78	197 Au <small>gold</small> 79	201 Hg <small>mercury</small> 80	204 Tl <small>thallium</small> 81	207 Pb <small>lead</small> 82	209 Bi <small>bismuth</small> 83	[209] Po <small>polonium</small> 84	[210] At <small>astatine</small> 85	[222] Rn <small>radon</small> 86											
[223] Fr <small>francium</small> 87	[226] Ra <small>radium</small> 88	[227] Ac* <small>actinium</small> 89	[261] Rf <small>rutherfordium</small> 104	[262] Db <small>dubnium</small> 105	[266] Sg <small>seaborgium</small> 106	[264] Bh <small>bohrium</small> 107	[277] Hs <small>hassium</small> 108	[268] Mt <small>meitnerium</small> 109	[271] Ds <small>darmstadtium</small> 110	[272] Rg <small>roentgenium</small> 111	Elements with atomic numbers 112-116 have been reported but not fully authenticated																	

* The lanthanoids (atomic numbers 58-71) and the actinoids (atomic numbers 90-103) have been omitted.

GENERAL CERTIFICATE OF SECONDARY EDUCATION

TWENTY FIRST CENTURY SCIENCE

CHEMISTRY A

A172/02

Unit A172: Modules C4, C5, C6 (Higher Tier)

MARK SCHEME

MAXIMUM MARK 60

Guidance for Examiners

Additional guidance within any mark scheme takes precedence over the following guidance.

1. Mark strictly to the mark scheme.
2. Make no deductions for wrong work after an acceptable answer unless the mark scheme says otherwise.
3. Accept any clear, unambiguous response which is correct, eg mis-spellings if phonetically correct (but check additional guidance).
4. Abbreviations, annotations and conventions used in the detailed mark scheme:

/	= alternative and acceptable answers for the same marking point
(1)	= separates marking points
not/reject	= answers which are not worthy of credit
ignore	= statements which are irrelevant - applies to neutral answers
allow/accept	= answers that can be accepted
(words)	= words which are not essential to gain credit
<u>words</u>	= underlined words must be present in answer to score a mark
ecf	= error carried forward
AW/owtte	= alternative wording
ORA	= or reverse argument

Eg mark scheme shows 'work done in lifting / (change in) gravitational potential energy' (1)

work done = 0 marks
work done lifting = 1 mark
change in potential energy = 0 marks
gravitational potential energy = 1 mark

5. Annotations:
The following annotations are available on SCORIS.

✓	= correct response
✗	= incorrect response
bod	= benefit of the doubt
nbod	= benefit of the doubt not given
ECF	= error carried forward
^	= information omitted
I	= ignore
R	= reject

6. If a candidate alters his/her response, examiners should accept the alteration.

7. Crossed out answers should be considered only if no other response has been made. When marking crossed out responses, accept correct answers which are clear and unambiguous.

Eg

For a one mark question, where ticks in boxes 3 and 4 are required for the mark:

Put ticks (✓) in the two correct boxes.

<input type="checkbox"/>
<input type="checkbox"/>
<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input type="checkbox"/>

This would be worth 0 marks.

Put ticks (✓) in the two correct boxes.

<input type="checkbox"/>
<input type="checkbox"/>
<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input type="checkbox"/>

This would be worth one mark.

Put ticks (✓) in the two correct boxes.

<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input checked="" type="checkbox"/>
<input type="checkbox"/>

This would be worth one mark.

8. The list principle:
If a list of responses greater than the number requested is given, work through the list from the beginning. Award one mark for each correct response, ignore any neutral response, and deduct one mark for any incorrect response, eg one which has an error of science. If the number of incorrect responses is equal to or greater than the number of correct responses, no marks are awarded. A neutral response is correct but irrelevant to the question.
9. Marking method for tick boxes:
Always check the additional guidance.

If there is a set of boxes, some of which should be ticked and others left empty, then judge the entire set of boxes.

If there is at least one tick, ignore crosses. If there are no ticks, accept clear, unambiguous indications, eg shading or crosses.

Credit should be given for each box correctly ticked. If more boxes are ticked than there are correct answers, then deduct one mark for each additional tick. Candidates cannot score less than zero marks.

Eg If a question requires candidates to identify a city in England, then in the boxes

Edinburgh	
Manchester	
Paris	
Southampton	


the second and fourth boxes should have ticks (or other clear indication of choice) and the first and third should be blank (or have indication of choice crossed out).

Edinburgh			✓			✓	✓	✓	✓	
Manchester	✓	x	✓	✓	✓				✓	
Paris				✓	✓		✓	✓	✓	
Southampton	✓	x		✓		✓	✓		✓	
Score:	2	2	1	1	1	1	0	0	0	NR

10. Three questions in this paper are marked using a Level of Response (LoR) mark scheme with embedded assessment of the Quality of Written Communication (QWC). When marking with a Level of Response mark scheme:
- Read the question in the question paper, and then the list of relevant points in the 'Additional guidance' column of the mark scheme, to familiarise yourself with the expected science. The relevant points are not to be taken as marking points, but as a summary of the relevant science from the specification.
 - Read the level descriptors in the 'Expected answers' column of the mark scheme, starting with Level 3 and working down, to familiarise yourself with the expected levels of response.
 - *For a general correlation between quality of science and QWC:* determine the level based upon which level descriptor best describes the answer; you may award either the higher or lower mark within the level depending on the quality of the science and/or the QWC.
 - *For high-level science but very poor QWC:* the candidate will be limited to Level 2 by the bad QWC no matter how good the science is; if the QWC is so bad that it prevents communication of the science the candidate cannot score above Level 1.
 - *For very poor or totally irrelevant science but perfect QWC:* credit cannot be awarded for QWC alone, no matter how perfect it is; if the science is very poor the candidate will be limited to Level 1; if there is insufficient or no relevant science the answer will be Level 0.

Question		Expected answers	Mark	Additional guidance
1		<p><i>metals:</i> the conclusion is supported but the melting point of A/ is only slightly higher than Mg</p> <p><i>non-metals:</i> the conclusion is supported for S, Cl and Ar but P is lower than S / P does not follow the trend</p> <p>could collect data for other periods to help evaluate the conclusions</p>	[3]	allow S not following trend
Total			[3]	

Question		Expected answers	Mark	Additional guidance
2		<p>the similarity is that they will both have lines</p> <p>the difference is that the lines will be different colours / the lines will be in different places / the lines will be in a different pattern</p>	[2]	
Total			[2]	


Question		Expected answers	Mark	Additional guidance
3	(a) 	<p>[Level 3] Most of the properties are discussed with clear comparison made. Discusses nuclear content, electron configuration and ions using actual numbers of particles /shells. All information in the answer is relevant, clear, organised and presented in a structured and coherent format. Specialist terms are used appropriately. Few, if any, errors in grammar, punctuation and spelling. (5-6 marks)</p> <p>[Level 2] Compares structure of nucleus, electronic structure and ions. For the most part the information is relevant and presented in a structured and coherent format. Specialist terms are used for the most part appropriately. There are occasional errors in grammar, punctuation and spelling. (3-4 marks)</p> <p>[Level 1] Some structure described but clear comparison may not be made. Discusses structure of nucleus and / or electronic configuration. Answer may be simplistic. There may be limited use of specialist terms. Errors of grammar, punctuation and spelling prevent communication of the science. (1-2 marks)</p> <p>[Level 0] Insufficient or irrelevant science. Answer not worthy of credit. (0 marks)</p>	[6]	<p>relevant points include:</p> <ul style="list-style-type: none"> • sodium has one electron in the outer shell, fluorine has 7 • sodium has 3 electron shells, fluorine has 2 • sodium has 11 protons, fluorine has 9 • sodium has 12 neutrons, fluorine has 10 • when sodium forms an ion it loses an electron • when fluorine forms an ion it gains an electron • both ions have a stable electron arrangement / full outer shell <p>for Level 2, accept 'different numbers of protons / neutrons / electron shells'</p>
		Total	[6]	

Question		Expected answers	Mark	Additional guidance
4	(a)	<u>orange gas</u> at start and <u>white solid</u> at end reaction takes 8-12 s / slower than iodine but faster than chlorine	[2]	
	(b)	sodium iodide NaI	[1]	both required for the mark, in either order
Total			[3]	

Question		Expected answers	Mark	Additional guidance															
5	(a)	<table style="width: 100%; border: none;"> <thead> <tr> <th></th> <th style="text-align: center;">true</th> <th style="text-align: center;">false</th> </tr> </thead> <tbody> <tr> <td>... reacts with cold water.</td> <td style="text-align: center;"><input checked="" type="checkbox"/></td> <td style="text-align: center;"><input type="checkbox"/></td> </tr> <tr> <td>... to form compounds.</td> <td style="text-align: center;"><input type="checkbox"/></td> <td style="text-align: center;"><input checked="" type="checkbox"/></td> </tr> <tr> <td>... more quickly than potassium.</td> <td style="text-align: center;"><input type="checkbox"/></td> <td style="text-align: center;"><input checked="" type="checkbox"/></td> </tr> <tr> <td>... is very unstable.</td> <td style="text-align: center;"><input type="checkbox"/></td> <td style="text-align: center;"><input type="checkbox"/></td> </tr> </tbody> </table>		true	false	... reacts with cold water.	<input checked="" type="checkbox"/>	<input type="checkbox"/>	... to form compounds.	<input type="checkbox"/>	<input checked="" type="checkbox"/>	... more quickly than potassium.	<input type="checkbox"/>	<input checked="" type="checkbox"/>	... is very unstable.	<input type="checkbox"/>	<input type="checkbox"/>	[2]	all 4 correct = 2 2 or 3 correct = 1 1 or 0 correct = 0 accept other indications of choice (eg lines or crosses)
	true	false																	
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... more quickly than potassium.	<input type="checkbox"/>	<input checked="" type="checkbox"/>																	
... is very unstable.	<input type="checkbox"/>	<input type="checkbox"/>																	
	(b)	Li and O ₂ as reactants correct balancing	[2]	the completed equation will be: 4Li + O ₂ → 2Li ₂ O															
Total			[4]																

Question		Expected answers	Mark	Additional guidance																														
6	(a)	<table border="1"> <thead> <tr> <th></th> <th>Andi</th> <th>Bea</th> <th>Carl</th> <th>Di</th> <th>none</th> </tr> </thead> <tbody> <tr> <td>...reproducibility</td> <td>✓</td> <td></td> <td></td> <td></td> <td></td> </tr> <tr> <td>...hypothesis</td> <td></td> <td></td> <td></td> <td>✓</td> <td></td> </tr> <tr> <td>...outlier</td> <td></td> <td></td> <td></td> <td></td> <td>✓</td> </tr> <tr> <td>...best estimate</td> <td></td> <td>✓</td> <td></td> <td></td> <td></td> </tr> </tbody> </table>		Andi	Bea	Carl	Di	none	...reproducibility	✓					...hypothesis				✓		...outlier					✓	...best estimate		✓				[2]	all four correct = 2 3 or 2 correct = 1 1 or 0 correct = 0
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...reproducibility	✓																																	
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...outlier					✓																													
...best estimate		✓																																
	(b)	C E B	[1]																															
Total			[3]																															

Question		Expected answers	Marks	Additional guidance
7		water, H ₂ O carbon dioxide, CO ₂	[2]	credit 1 mark for any two names/formulae correct
Total			[2]	

Question	Expected answers	Mark	Additional guidance
8 	<p>[Level 3] Chooses aluminium and uses its properties to explain suitability. Uses properties of other metals to explain their lack of suitability. Refers to compromise of properties for purpose. All information in the answer is relevant, clear, organised and presented in a structured and coherent format. Specialist terms are used appropriately. Few, if any, errors in grammar, punctuation and spelling. (5-6 marks)</p> <p>[Level 2] Chooses aluminium and uses its properties to explain suitability. Makes some reference to properties of other metals but does not explain their lack of suitability. For the most part the information is relevant and presented in a structured and coherent format. Specialist terms are used for the most part appropriately. There are occasional errors in grammar, punctuation and spelling. (3-4 marks)</p> <p>[Level 1] Chooses a metal other than aluminium. Makes some relevant comments about its suitability. Answer may be simplistic. There may be limited use of specialist terms. Errors of grammar, punctuation and spelling prevent communication of the science. (1-2 marks)</p> <p>[Level 0] Insufficient or irrelevant science. Answer not worthy of credit. (0 marks)</p>	[6]	relevant points include: <ul style="list-style-type: none"> • aluminium has lowest density so cable can span long distances • aluminium has good resistance to corrosion so cables will last a long time • aluminium has reasonable conductivity but this is a compromise • aluminium is quite cheap so lots of cables can be used for reasonable cost • gold has very good conductivity but is too heavy and is too expensive • iron is cheap but is too heavy and corrodes too easily • copper has good conductivity but is too heavy and too expensive
Total		[6]	

Question			Expected answers	Mark	Additional guidance
9	(a)	(i)	the (trend shown in the) graph suggests that use of/demand for copper will continue to increase however, large amounts of ore need to be mined to recover a small percentage of copper so we may not be able to mine enough ore to meet the demand / only a small proportion of the copper in the Earth's crust is close enough to the surface to be mined economically	[2]	for full marks the explanation must be expressed in a logical coherent order
		(ii)	used copper could be recycled	[1]	
	(b)	(i)	$\frac{63.5 + 63.5}{63.5 + 63.5 + 16} \times 40$ = 35.5	[2]	
		(ii)	heated reduced oxygen oxidised oxygen	[1]	all correct for the mark reject "reacted" / "mixed" / "treated" etc.
Total				[6]	


Question		Expected answers	Mark	Additional guidance
10	(a)	lead and bromine form lead at the negative electrode and bromine at the positive electrode	[2]	
	(b)	e ⁻ and Na	[1]	both required for the mark
Total			[3]	

Question		Expected answers	Mark	Additional guidance
11		<input type="checkbox"/> Electrons are shared between atoms. <input checked="" type="checkbox"/> The nucleus of each bonded atom ... <input checked="" type="checkbox"/> <input type="checkbox"/> <input type="checkbox"/>	[2]	
Total			[2]	

Question		Expected answers	Marks	Additional guidance
12	(a)	citric acid	[1]	
	(b)	the acid and alkali need to dissolve in water to produce H ⁺ (aq) and OH ⁻ (aq) ions (which are free to react)	[1]	
Total			[2]	

Question	Expected answers	Marks	Additional guidance
13	<p>she should add acid in small amounts near the end point so that she does not add more acid than necessary to neutralise the alkali</p> <p>OR</p> <p>she should swirl the flask between each addition of acid so that the acid mixes completely before adding any more</p> <p>OR</p> <p>she should look carefully for first (permanent) colour change so that she does not add more acid than necessary to neutralise the alkali</p>	[2]	<p>ignore "do it (more) carefully"</p> <p>for full marks the action Mary takes should be coherently linked to the resulting improvement in accuracy</p>
Total		[2]	




Question		Expected answers	Marks	Additional guidance
14	(a)	gram formula mass of $\text{MgSO}_4 = 24 + 32 + 64 = 120 \text{ g}$ gram formula mass of $\text{Mg} = 24 \text{ g}$ $\frac{3}{24} \times 120 = 15$	[2]	
	(b)	$20 \div 10 = 2 \text{ cm}^3/\text{s}$	[1]	correct working, answer and units required for the mark
	(c)	experiment A because a larger mass of magnesium pieces will give a higher rate of reaction, so more gas will have been produced by 10s and a larger mass of reactant will produce a greater volume of product/gas/hydrogen	[3]	for full marks the explanation must be expressed in a logical and coherent order

Question		Expected guidance	Gd	Additional guidance
14	(d) 	<p>[Level 3] Answer demonstrates an understanding of the nature of the particles involved and the effect of their collisions on the rate of reaction. All information in the answer is relevant, clear, organised and presented in a structured and coherent format. Specialist terms are used appropriately. Few, if any, errors in grammar, punctuation and spelling. (5-6 marks)</p> <p>[Level 2] Answer deals with one aspect, eg collision frequency, but does not discuss the nature of the colliding species. For the most part the information is relevant and presented in a structured and coherent format. Specialist terms are used for the most part appropriately. There are occasional errors in grammar, punctuation and spelling. (3-4 marks)</p> <p>[Level 1] Answer shows an awareness of the basic premise, that of collisions, but has difficulty identifying the reacting species and sees the reaction in terms of number of collisions rather than frequency. Detail of what constitutes a low-level answer. Answer may be simplistic. There may be limited use of specialist terms. Errors of grammar, punctuation and spelling prevent communication of the science. (1-2 marks)</p> <p>[Level 0] Insufficient or irrelevant science. Answer not worthy of credit. (0 marks)</p>	[6]	<p>relevant points include:</p> <ul style="list-style-type: none"> hydrogen/H⁺ ions from the acid react with magnesium atoms at the surface of the magnesium using smaller pieces of magnesium gives a larger surface area allowing the hydrogen/H⁺ ions to collide more frequently with the magnesium atoms which will increase the rate of reaction <p>reject references to increased speed of movement reject references to increased concentration of the acid</p>
		Total	[12]	

Question		Expected answers	Marks	Additional guidance
15	(a)	HCl in reactant box H ₂ O and CO ₂ in product boxes, in either order equation correctly balanced	[3]	the completed equation will be: $\text{CuCO}_3 + 2\text{HCl} \rightarrow \text{CuCl}_2 + \text{H}_2\text{O} + \text{CO}_2$
	(b)	exothermic	[1]	
Total			[4]	

Assessment Objectives (AO) Grid

(includes quality of written communication )

Question	AO1	AO2	AO3	Total
1			3	3
2	1	1		2
3 	3	3		6
4(a)	1	1		2
4(b)	1			1
5(a)	2			2
5(b)	1	1		2
6(a)		2		2
6(b)	1			1
7	1	1		2
8 		3	3	6
9(a)(i)			2	2
9(a)(ii)		1		1
9(b)(i)		2		2
9(b)(ii)	1			1
10(a)		2		2
10(b)	1			1
11	2			2
12(a)	1			1
12(b)		1		1
13	2			2
14(a)		2		2
14(b)		1		1
14(c)		2	1	3
14(d) 	4	2		6
15(a)	2	1		3
15(b)	1			1
Totals	25	26	9	60